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6.2 Calculate the heat energy  $q$  in kJ required to heat 2.500 g of water from 25.0 °C to 60.0 °C.  $c = 4.18 \text{ J g}^{-1} \text{ K}^{-1}$  and  $\Delta T = 0.60 \text{ K}$ . A. 6.6 B. 6.3 C. 6.27 D. 6.270

6.3 Which molecule has an index of hydrogen deficiency (IHD) = 1? A.  $\text{CH}_4$  B.  $\text{CCl}_4$  C.  $\text{C}_6\text{H}_6$  D.  $\text{C}_6\text{H}_8$

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39.10 37 Rb 85.47 55 Cs 132.91 87 Fr  
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atomic mass 25 Mn 54.94 43 Tc 98.91  
75 Re 186.21 61 Pm 146.92 93 Np  
(237) 26 Fe 55.85 44 Ru 101.07 76 Os  
190.21 62 Sm 150.35 94 Pu (242) 77 I  
63 Eu 151.96 95 Am (243) 27 Co  
58.93 45 Rh 102.91 r 192.22 28 Ni  
58.71 46 Pd 106.42 78 Pt ...

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question (the volume one) on Topic 1, the shape of that weird molecule on Topic 4, and energetic manipulations on Topic 5. But overall, this paper was okay though. Past papers really helped. P.S. Topic 2 question on the number of electrons in NUCLEUS was LIT. Can you imagine how the IB

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